Version No.			ROLL NUMBER						STATE AND SECTION OF THE SECTION OF		
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(1)	(1)	(1)	(1)	(1		(1)	(1)	(1)	(1)	(1)	SLAMABADINA
(2)	2	2	2	(2	_	2	2	2	2	2	
(3)	3	(3)	(3)	(3		3	3	3	(3)	3	Answer Sheet No
(4)	4	4	<ul><li>(4)</li></ul>	(4		4	4	4	4	4	
(5)	(5)	(5)	(5)	(5		(5)	(5)	(5)	(5)	(5)	Sign. of Candidate
6	(6)	6	<u>(6)</u>	(6		6	6	6	6	6	
(7)	7	7	(7)	(7		7	7	7	(7)	(7)	
(8)	(8)	(8)	(8)	(8		(8)	(8)	(8)	(8)	(8)	Sign. of Invigilator
9	9	(9)	9	(9		9	(9)	(9)	(9)	9	
						CHI					7 1
						SEC					
						SEC Time			`		
			-	•	-						be answered on this page and hande not allowed. <b>Do not use lead pencil.</b>
.1	Fil	l the	releva	ant bul	ble fo	r eac	ch pa	rt. E	Cach	part	carries one mark.
	(1)										Il be formed by an element of group $2p^6 3s^2$ ?
			HA na A.	A <sup>+ 3</sup>	ection	ic coi	inigu	Tauo	н 18 В		A <sup>+2</sup>
			C.	$A^{+1}$					D		$A^{-2}$
	(2)			one of er pairs				airs o	of sul	oshel	l has the lowest energy as compared
			A.	1s,2s					В		2s,2p
			C.	3s, 3p					D		3s, 4s
	(3)					llow	ing Is	sotop			in nuclear reactors?
			A. C.	U-234 U-235					B D		U-238 U-233
	(4)		How n A.	nany m 8 x 6.			oxyg	gen g	as co B		is one mole of oxygen gas?
			C.		0.022  x				ע	•	$6.022 \times 10^{23}$
				32 X C	· -				D		$6.022 \times 10^{23}$ $16 \times 6.022 \times 10^{23}$
	(5)		The ve					tant i			$16 \times 6.022 \times 10^{23}$
	(5)		The va A.	riable	that is	kept		tant i		arles	
	(5)				that is erature	kept		tant i	n Ch	arles	16 x 6.022 x 10 <sup>23</sup> ' Law is:
			A. C.	Temp Pressu	that is erature ire	kept e	cons		n Ch B D	arles	16 x 6.022 x 10 <sup>23</sup> 'Law is: Volume Volume & Temperature
	<ul><li>(5)</li><li>(6)</li></ul>		A. C.	riable Temp	that is erature ire	kept e	cons		n Ch B D	arles ollow	16 x 6.022 x 10 <sup>23</sup> 'Law is: Volume Volume & Temperature

(7)	Press	ure Cooker works on the princ	ciple of	relationship of boiling point with:									
	A.	External Pressure	B.	Evaporation									
	C.	Boyle's law	D.	Volume									
(8)	17g o	f NH <sub>3</sub> is dissolved in 1 dm <sup>3</sup> or	f solutic	on, its molarity will be:									
(0)	A.	1	В.	2									
	C.	3	D.	4									
(9)	In Ha	S, the oxidation state of Sulph	ur ic.										
(9)	A.	+1	B.	+ 2									
	C.	· <del>-</del>		+ 2 -2									
	C.	- 1	D.	-2									
(10)	The compound having Hydrogen bonding among its molecule is:												
	A.	$C_6H_6$	B.	MgO									
	C.	CH <sub>4</sub>	D.	$H_2O$									
(11)	Metallic Character increases down the group, which one of the following is the												
	most	metallic:											
	A.	Rb	B.	Cs									
	C.	Na	D.	K									
(12)	The most electronegative element in the group VIIA is:												
	A.	F	B.	Cl									
	C.	Br	D.	I									



xi.

# Federal Board SSC-I Examination Chemistry Model Question Paper (Curriculum 2006)

Time allowed: 2.40 hours Total Marks: 53

Note: Answer all parts from Section 'B' and all questions from Section 'C' on the E-sheet. Write your answers on the allotted/given spaces.

**SECTION – B** (Marks 33) **Q.2** Attempt all parts from the following. All parts carry equal marks.  $(11 \times 3 = 33)$ Calculate the number of molecules in 4.5 moles of Carbon dioxide. (1+2)Calculate the mass of one Hydrogen atom in gram. (1+2)ii. Draw Bohr's Atomic Model for Potassium 19K<sup>39</sup> indicating the location of electrons, protons and neutrons. (1+1+1)State Charles's Law. Derive its mathematical expression. iii. (1+2)Define ionic bond. Give one example of two elements forming an ionic bond iv. between them. (1+2)Write two similarities and two differences between isotopes. (1.5+1.5)v. vi. Elements are unstable in free state except noble gases. Explain how elements attain stability? (1+2)Why is an atom always electrically neutral? Give reason (1+2)vii. Write electronic configuration of Aluminum 13Al<sup>27</sup>. Identify its group and period. OR How does the change in temperature affect the Vapour Pressure of a liquid? Show with the help of graph. (1+2)viii. How will you prepare 250 cm<sup>3</sup> of 0.025M Na<sub>2</sub>SO<sub>4</sub> solution from a stock solution of 2M Na<sub>2</sub>SO<sub>4</sub>? (1+2)Identify the oxidizing and reducing agents in the following reaction with indicating oxidation number: (1.5+1.5)→ 2HCl + S  $H_2S + Cl_2$ a.  $\longrightarrow$  MgCl<sub>2</sub> + H<sub>2</sub> Mg + 2HCl -Define corrosion. How is corrosion prevented by cathodic protection? (1+2)Enlist the name of three noble metals? (1+1+1)OR Why is the boiling point of water at the top of Mount Everest  $70^{\circ}$ C. Give a reason? (1+2)

(1.5+1.5)

Discuss why sugar is soluble in water but petrol is not?

## **SECTION – C** (Marks 20)

Note: Attempt all questions. Marks of each question are given within brackets.

Q.3 What are the type of bonds responsible for the formation of  $F_2$ ,  $O_2$  and  $N_2$ ? Explain the formation of bond with the help of structures. (2+2+2)

#### OR

Describe Rutherford's Experiment with diagram and its conclusions. (3+3)

Q.4 Describe three importance of intermolecular forces in our life. (2+2+2)

#### OR

Describe the trend of Ionization Energy in the Period and group. Explain with reasons. (3+3)

Q.5 Explain the working and construction of Daniel Cell with the help of a labelled diagram. (2+2)

### OR

Identify the relationship between electronic configuration and the position of an element in the periodic table.  $_{35}Br^{70}$  and  $_8O^{16}$  (2+2)

Q.6 By using following reactions. Discuss the reactivity (2+2)

i)  $2KI + Br_2 \longrightarrow 2KBr + I_2$ 

ii)  $KBr + Cl_2 \longrightarrow 2KCl + Br_2$ 

\* \* \* \* \*

					SUPLI	EMENT	TARY 1	ABLE						
Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14
Symbol	Н	He	Li	Ве	В	С	N	0	F	Ne	Na	Mg	Al	Si
Mass no	1	4	7	9	11	12	14	16	19	20	23	24	27	28
Atomic No	15	2	16	17	18	19	20	31	32	33	34	35	36	37
Symbol	Р	He	S	CI	Ar	K	Ca	Ga	Ge	As	Se	Br	Kr	Rb
Mass no	31	4	32	35	40	39	40	70	73	74	79	80	84	85
Atomic No	38	49	50	51	52	53	54	55	56	81	82	83	84	85
Symbol	Sr	In	Sn	Sb	Те	ı	Xe	Cs	Ва	TI	Pb	Bi	Ро	At
Mass no	88	115	119	122	128	127	131	133	137	204	207	208	209	210