

Version No.			

ROLL NUMBER						

0 0 0 0

1 1 1 1

2 2 2 2

3 3 3 3

4 4 4 4

5 5 5 5

6 6 6 6

7 7 7 7

8 8 8 8

9 9 9 9

0 0 0 0 0 0 0

1 1 1 1 1 1 1

2 2 2 2 2 2 2

3 3 3 3 3 3 3

4 4 4 4 4 4 4

5 5 5 5 5 5 5

6 6 6 6 6 6 6

7 7 7 7 7 7 7

8 8 8 8 8 8 8

9 9 9 9 9 9 9

Answer Sheet No. _____

Sign. of Candidate _____

Sign. of Invigilator _____

CHEMISTRY HSSC-I (3rd Set)

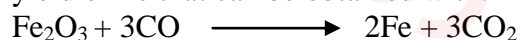
SECTION – A (Marks 17)

Time allowed: 25 Minutes

Section – A is compulsory. All parts of this section are to be answered on this page and handed over to the Centre Superintendent. Deleting/overwriting is not allowed. **Do not use lead pencil.**

Q.1 Fill the relevant bubble for each part. Each part carries one mark.

- (1) When **6 MOLES** of CO are used in the following reaction, predict the maximum yield of Fe that can be obtained with:



- A. 1 mole of Fe_2O_3 B. 2 mole of Fe_2O_3
 C. 3 mole of Fe_2O_3 D. 4 mole of Fe_2O_3

- (2) Quote what Mosely was able to determine with the discovery of **X-Rays**:

- A. The number of protons in target metal atom
 B. The number of neutrons in target metal atom
 C. The mass number of target metal atom
 D. The atomic mass of target metal atom

- (3) Categorize the polar molecule among the following:

- A. $\begin{array}{c} \text{Cl} \\ | \\ \text{Cl}-\text{C}-\text{Cl} \\ | \\ \text{Cl} \end{array}$ B. $\begin{array}{c} \text{H} \\ | \\ \text{H}-\text{C}-\text{H} \\ | \\ \text{H} \end{array}$
 C. $\begin{array}{c} \text{H} \\ | \\ \text{H}-\ddot{\text{N}}-\text{H} \\ | \\ \text{H} \end{array}$ D. $\text{O}=\text{C}=\text{O}$

- (4) What are the conditions at which Real gases behave ideally:

- A. Standard temperature and pressure
 B. High pressure and low temperature
 C. Low pressure and high temperature
 D. Low temperature and high pressure

- (5) Gases exert pressure on the walls of their container. Predict it is due to:
- A. Force of attraction between their molecules
- B. Elastic collisions between their molecules
- C. Vander Waal's forces between their molecules
- D. Force of repulsion between their molecules
- (6) Label which of the following valence shell electronic configurations are correct **EXCEPT**:
- A. $\begin{array}{cccc} \uparrow\downarrow & \uparrow\downarrow & \uparrow & \uparrow \\ 2s & 2px & 2py & 2pz \end{array}$
- B. $\begin{array}{cccc} \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ 2s & 2px & 2py & 2pz \end{array}$
- C. $\begin{array}{cccc} \uparrow\downarrow & \uparrow\downarrow & \uparrow & \uparrow \\ 2s & 2px & 2py & 2pz \end{array}$
- D. $\begin{array}{cccc} \uparrow\downarrow & \uparrow\downarrow & & \\ 2s & 2px & 2py & 2pz \end{array}$
- (7) Light emitted or absorbed in form of energy packets called photons. What is the energy of photon?
- A. $E=hC$
- B. $E=hc/\lambda$
- C. $E=h\lambda$
- D. $E=c/\lambda$
- (8) Predict which one has highest boiling point among the following:
- A. Helium (He)
- B. Neon (Ne)
- C. Argon (Ar)
- D. Krypton (Kr)
- (9) Identify the temperature at which water (H₂O) exhibits lower density:
- A. 34°C
- B. 24°C
- C. 14°C
- D. 4°C
- (10) Analyze what is true for the following chemical reaction?
- $$\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$$
- A. $k_p > k_c$
- B. $k_n < k_c$
- C. $k_p = k_n$
- D. $k_p > k_n$
- (11) An acid X considered to be stronger one, when the value of its Dissociation constant (K_a) is:
- A. 1.8×10^{-16}
- B. 1.0×10^{-4}
- C. 1.0×10^9
- D. 1.0×10^{10}
- (12) For a chemical reaction A the rate equation is given below:
 $\text{Rate} = k[\text{O}_3]^2[\text{O}_2]^{-1}$
 Distinguish the **INCORRECT** statement about the reaction A:
- A. The rate of reaction increases four time when $[\text{O}_3]$ is doubled and $[\text{O}_2]$ is kept constant.
- B. The rate of reaction increases twice when $[\text{O}_2]$ is doubled and $[\text{O}_3]$ is kept constant
- C. The rate of reaction becomes half when $[\text{O}_2]$ is doubled and $[\text{O}_3]$ is kept constant.
- D. The overall order of this reaction A is First order.
- (13) Identify the temperature at which the solubility of CuSO₄ in water is maximum at:
- A. 80°C
- B. 90°C
- C. 100°C
- D. 110°C

- (14) If the following reaction is performed at STP, predict which one of the following statement is correct for this reaction?
- $$\text{CO}_2(\text{g}) \longrightarrow \text{C}_{(\text{s})} + \text{O}_{2(\text{g})}$$
- A. ΔH is positive and the temperature of surrounding is increased.
- B. ΔH is positive and the temperature of surrounding is decreased.
- C. ΔH is negative and the temperature of surrounding is increased.
- D. ΔH is negative and the temperature of surrounding is decreased.
- (15) In the following reaction, quote the correct statement:
- $$\text{Ca}(\text{s}) + \text{Cl}_2(\text{g}) \longrightarrow \text{CaCl}_2$$
- A. Ca is oxidizing agent B. Cl_2 is reducing agent
- C. Ca is reducing agent D. CaCl_2 is reducing agent
- (16) Consider the following general chemical reaction at equilibrium. If it's $K_c = 1$, what is the best statement for the chemical reaction.
- $$\text{Reactant} \rightleftharpoons \text{Product}$$
- A. Concentration of reactant = Concentration of product
- B. Concentration of reactant > Concentration of product
- C. Rate of forward reaction > Rate of backward reaction
- D. Rate of forward reaction < Rate of backward reaction
- (17) When two atoms bond with each other through overlapping of their sp^3 hybrid orbitals. Predict which one of the following statements best describes the nature of the bond between these atoms?
- A. This is a Pi (π) covalent bond
- B. This is a Sigma (σ) covalent bond
- C. This is ionic bond
- D. This is hydrogen bond.
-

CHEMISTRY HSSC-I (3rd Set Solution)

SECTION – A (Marks 12)

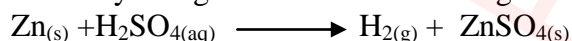
Time allowed: 20 Minutes

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	15	17
B	A	C	C	B	D	B	D	A	C	D	B	A	B	C	A	B

SECTION – B

Q.2

- i. 8.7×10^{24} H_2 molecules are produced by the reaction of zinc and 30% W/W H_2SO_4 solution of density 1.25 g/cm^3 . For the following reaction:



Calculate the volume of sulphuric acid solution used.?

(Zn = 65, O = 16, S = 32, H = 1 g/mol)

Answer:

It this reaction 30% W/W H_2SO_4 is used, it means 30 g of sulfuric acid is used to make 100 g of solution.

Molecules of H_2 produced = 8.7×10^{24}

Moles of H_2 = $8.7 \times 10^{24} / 6.022 \times 10^{23}$

Moles of H_2 = 14.45

According to chemical equation:

1 mole of H_2 is produced from: 1 mole of H_2SO_4

14.45 moles of H_2 is produced from : 14.45 mole of H_2SO_4

Mass of H_2SO_4 = moles x molar mass

Mass of H_2SO_4 = 14.45×98

Mass of H_2SO_4 = 1416.28 g

We can calculate the volume of solution for 30% W/W solution:

Mass of solution: 4720.93 g

$d = m/v$

By putting the values:

$1.25 \text{ g/cm}^3 = 4720.93 \text{ g} / v$

$V = 3776.74 \text{ cm}^3$

- ii. Briefly describe the miscibility of phenol and water system.

Answer:

When equal volumes of phenol and water are mixed into each other, two layers are formed. The lower layer consists of small amount of water dissolved in phenol, while the upper layer

consists of small amount of phenol dissolved in water. It is observed that, upper layer is 5% solution of phenol in water and lower layer is 30% water in phenol at 25 °C.

So, we conclude that phenol and water are partially miscible liquids in each other. Water is a polar molecule and phenol has –OH group, which is hydrophilic and non-polar benzene ring, which is hydrophobic. The –OH group can interact with water by hydrogen bonding and dissolve but benzene does not. Because of this reason, the phenol water system is a partially miscible solution.

iii. Describe briefly Zeeman and Stark effect on Bohr atomic model.

Answer:

Zeeman effect is the splitting of the line of a spectrum when the source of the spectrum is exposed to a magnetic field. Whereas, Stark effect is the splitting of spectral lines observed when the radiating atoms, ions, or molecules are subjected to a strong electric field. The electric analogue of the Zeeman effect (i.e., the magnetic splitting of spectral lines). For example, according to Azimuthal quantum number of value of $l=0$ there is no effect of magnetic as well as electric field. If $l= 1$. Then it splits up into three spectral lines i.e. $+1, 0, -1$.

iv. List the quantum numbers value of 5p.

Answer:

The principal quantum number $n = 5$,

The azimuthal quantum number $l = 1$

The magnetic quantum number $m = +1, 0, -1$

The spin quantum number $+1/2$ and $-1/2$

v. Interpret the change in the bond energy in the following.

Bonds	C-F	C-Cl	C-Br	C-I
Energy KJ/mol	485	329	276	240

Answer:

The bond energy goes on decreasing from C-F to C-I. The bond energy of C-F bond is highest in the list. The reason is that the highest electronegativity and small size of F. The bond energy of C-Cl is lesser than C-F. The reason is that the low electronegativity and large size of Cl. Similarly, the bond energy is lowest in case of C-I. The reason is that the lowest electronegativity and largest size of I. Hence, in moving from C-F to C-I, electronegativity difference decreases and bond length increases, which will decrease the bond energy.

vi. Demonstrate the values of bond pair, lone pair and total electron pair in AsH_3 by VSEPR theory, and draw its structure.

Answer:

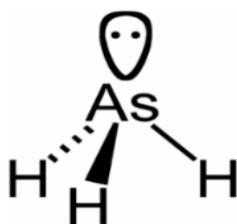
As has 5 valence electrons. This is AB_3E type structure according to VSEPR.

Total electron pairs = 4.

Bond pairs = 3

Lone pair = 1

Structure of AsH_3 according to VSEPR is trigonal pyramidal:



Lone pair bond pair repulsion is greater than bond pair, bond pair repulsion. Hence bond angle will be 107.5° . This is AB_3E type structure according to VSEPR.

vii. Show by derivation the unit of density in the expression $d = \frac{PM}{RT}$ when $R = 0.0821 \text{ atm dm}^3 \text{ mol}^{-1} \text{ K}^{-1}$.

Answer:

Units of density:

Putting the units of all parameters:

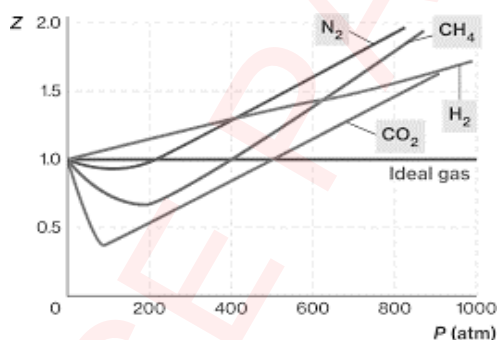
$$d = \frac{\text{atm} \cdot \text{g} \cdot \text{mol}^{-1}}{\text{atm} \cdot \text{dm}^3 \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot \text{K}}$$

$$d = \text{g} / \text{dm}^3$$

So the units of the density according to given equation is:

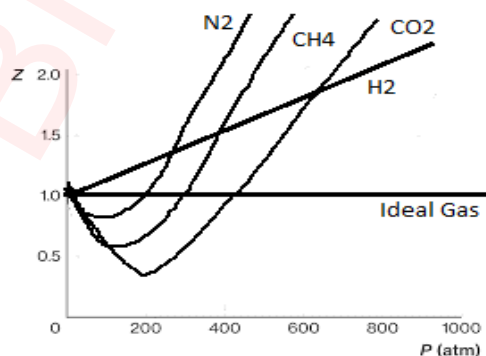
$$\text{g} / \text{dm}^3$$

viii. Deviation from ideal behavior is shown in following graph at 25°C . Sketch the deviation from ideal behavior at 5°C .



Answer:

Deviation from ideal behavior at 5°C :



ix. A gas collected over water at 10°C and 873 torr. If gas occupies 90cm^3 , calculate the volume of dry gas at STP when aqueous tension is 9.2 torr.

Answer:

Data:

$$V_1 = 90 \text{ cm}^3$$

From Dalton's law $P_g = P_t - P_{\text{H}_2\text{O}}$

$$P_1 = 873 \text{ torr} - 9.2 \text{ torr} = 863.8 \text{ torr}$$

$$T_1 = 283 \text{ K}$$

$$P_2 = 760 \text{ torr}$$

$$V_2 = ?$$

$$T_2 = 273 \text{ K}$$

According to the general gas equation

$$\frac{V_1 P_1}{T_1} = \frac{V_2 P_2}{T_2}$$

$$V_2 = \frac{V_1 P_1 T_2}{T_1 P_2}$$

$$V_2 = \frac{90 \text{ cm}^3 \times 863.8 \text{ torr} \times 273 \text{ K}}{283 \text{ K} \times 760 \text{ torr}}$$

$$V_2 = 98.7 \text{ cm}^3$$

x. Describe the surface tension of the given solvents in term of intermolecular forces.

Solvent	Surface Tension
Water (H_2O)	7.275×10^2
Hexane (C_6H_{14})	1.84×10^2

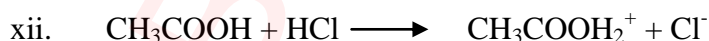
Answer:

Surface tension of the water is higher as compare to the surface tension of hexane. Water is a polar substance while hexane is a non-polar substance. Water molecules possess hydrogen bonding but in case of hexane, it possesses London dispersion force, it contains more number of carbon atoms i.e. 6, which enhance the LDF. As, LDF is weaker than hydrogen bonding, because of this reason, the surface tension of the water is higher as compared to hexane.

xi. Ice floats at the surface of the water. Tabulate its three advantages.

Answer:

1	Because of this phenomenon, the temperature of water in depth will be maintained.
2	The life under water can survive because of normal temperature.
3	Ice on the surface will not allow more heat to dissipate from the water below the surface.



What is conjugate acid and conjugate base pair, identify conjugate acid and conjugate base.

Answer: In the Bronsted-Lowry definition of acids and bases, a conjugate acid-base pair consists of two substances that differ only by the presence of a proton (H^+). A conjugate acid

is formed when a proton is added to a base, and a conjugate base is formed when a proton is removed from an acid.

Identification:

Conjugate acid = $\text{CH}_3\text{COOH}_2^+$

Conjugate base: Cl^-

xiii. Calculate pH of $1 \times 10^{-3} \text{ M Fe(OH)}_3$.

Answer:

Fe(OH)_3 is a base so first of all we will calculate the OH^- .



Upon ionization 3 OH^- ions are produced. So, concentration will become

$$3 \times 10^{-3} \text{ M}$$

$$\text{pOH} = -\log [\text{OH}^-]$$

$$\text{pOH} = -\log 3 \times 10^{-3}$$

$$\text{pOH} = 2.52$$

As we know that:

$$\text{pH} + \text{pOH} = 14$$

So

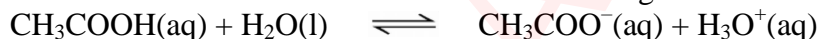
$$\text{pH} = 14 - 2.52$$

$$\text{pH} = 11.48$$

xiv. Briefly explain buffer action with the help of CH_3COOH and CH_3COONa .

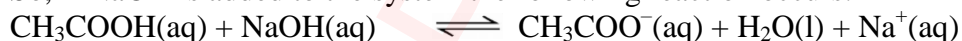
Answer:

A mixture of a weak acid and its salt with a strong base serves as an acidic buffer.



CH_3COOH is a weak acid, while CH_3COO^- is a strong conjugate base

So, if NaOH is added to the system the following reaction occurs:



The reaction shows that some acetic acid reacts with the OH^- from the base and converts it to water. Therefore, the pH will not change as drastically as it would have without the buffer. Sodium acetate ionizes completely providing the acetate ion, CH_3COO^- .

If HCl is added:



When HCl is added in the solution H^+ will react with acetate ion which results in the formation of acetic acid, thus maintaining the pH of the solution.

xv. Discuss briefly the colligative property of $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ and $\text{CO(NH}_2)_2$.

Answer:

Sucrose and urea both are non volatile and non electrolyte solutes. When 342 g $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ or 60 g $\text{CO(NH}_2)_2$ are dissolved in one kg water (a volatile solvent), there will be following observations in the solution:

1. Elevation of boiling point 100.52°C
2. Depression of freezing point -1.86°C

Hence colligative properties shown by sucrose and urea do not depend upon mass but depend on number of moles and molecules when they are dissolved in a volatile solvent.

xvi. Prove first law of thermodynamics and show that $\Delta E = q_v$.

Answer:

First law of thermodynamics:

"This law is also called the law of conservation of energy. This law is stated as, "energy of the universe is constant".

OR

"Energy can neither be created nor destroyed but can change from one form to another."

Mathematically.

$$\Delta E = q + W$$

$$\Delta E = q + W$$

Where ΔE = Internal energy changes q = Heat and W = work

Since $W = -P \Delta V$

At constant volume:

If the volume of gas does not change, no work is done, ($\Delta V = 0$). By applying the first law of thermodynamics.

$$\Delta E = q_v - P \Delta V$$

$$\Delta E = q_v - 0 \quad (\Delta V = 0)$$

$$\Delta E = q_v$$

So the increase of heat at constant volume (q_v) increases only the internal energy (ΔE) of the system and work done is zero.

xvii. Differentiate between exothermic and endothermic reactions with thermochemical equations.

Answer:

Exothermic	Endothermic
<p>Chemical reactions that release energy are called exothermic. In exothermic reactions, more energy is released when the bonds are formed in the products than is used to break the bonds in the reactants.</p> <p>Thermochemical equations:</p> $2 \text{NO}_2(\text{g}) \longrightarrow \text{N}_2\text{O}_4(\text{g}) \quad \Delta H = - 57.6 \text{ kJ}$ $\text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g}) \longrightarrow \text{PCl}_5(\text{g}) \quad \Delta H = - 92.5 \text{ kJ}$	<p>Chemical reactions that absorb energy are called endothermic reactions. In endothermic reactions, more energy is absorbed when the bonds in the reactants are broken than is released when new bonds are formed in the products.</p> <p>Thermochemical equations:</p> $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2 \text{NO}(\text{g}) \quad \Delta H = + 181 \text{ kJ}$ $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2 \text{NO}(\text{g}) \quad \Delta H = + 181 \text{ kJ}$

xviii. Briefly describe cleavage planes and habit of crystal.

Answer:

Cleavage plane:

A cleavage plane is the fracture of a crystal or metal by crack propagation across a crystallographic plane or cleavage plane, or the tendency to cleave or split along definite crystallographic planes.

Habit of crystal:

Crystal habit is the characteristic external shape of an individual crystal or groups of crystals. A single crystal's habit is a description of its general shape and its crystallographic forms, and how well developed each are.

xix. Complete the reaction on the bases of reduction potentials of $\text{Pb}^{+2}/\text{Pb} = -0.36\text{v}$ and $\text{Fe}^{+2}/\text{Fe} = -0.44\text{v}$
 $\text{PbSO}_4 + \text{Fe} \longrightarrow$

Answer:

Iron is ranked higher in activity series as compared to lead so it can displace lead.

$$\begin{aligned} E_{\text{cell}} &= E_{\text{cathode}} - E_{\text{anode}} \\ &= -0.36 - (-0.44) \\ &= +0.08\text{v} \end{aligned}$$

Feasible reaction is



xx. Calculate the energy of ${}_3\text{Li}^{+2}$ for Balmer series when electron jumps from $n=5$.

Answer:

In Balmer series electron jumps from $n_2 = 5$ to $n_1 = 2$.

So, we will use following formula:

$$\Delta E = \frac{mZ^2e^4}{8\epsilon_0^2h^2} \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$Z = 3$ for Li^{+2} .

$n_1 = 2$

$n_2 = 5$

$$\Delta E = 2.18 \times 10^{-18} \times (3)^2 (1/2^2 - 1/5^2)$$

$$\Delta E = 2.18 \times 10^{-18} \times (1/4 - 1/25)$$

$$\Delta E = 4.57 \times 10^{-19} \text{ J}$$

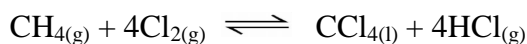
SECTION – C

Q.3 a. $\text{CH}_{4(g)} + 4\text{Cl}_{2(g)} \rightleftharpoons \text{CCl}_{4(l)} + 4\text{HCl}_{(g)}$
Derive Kc expression at equilibrium of given reaction.

(06)

Answer:

Kc expression:



Initial conc.	a/v	b/v	0
Eq. Conc.	a-x /v	b-4x/v	4x/v
$K_c = \frac{[\text{HCl}]^4}{[\text{CH}_4][\text{Cl}_2]^4}$			
$K_c = \frac{(4x)^4 / v^4}{(a-x)/v (b-4x)^4 / v^4}$			
$K_c = \frac{256(x)^4 v}{(a-x)(b-4x)^4}$			

b. Describe that increase in collision energy by increasing the temperature can improve the collision frequency. (3+2+2)

Answer:

According to Arrhenius equation:

$$k = A \cdot e^{-\frac{E_a}{RT}}$$

- k - rate constant
- A - Arrhenius constant
- E_a - activation energy
- R - gas constant
- T - temperature

According to above equation, A is frequency constant which explains the effect of change in temperature on collision frequency.

We know that a chemical system can be made up of atoms. These particles are moving around in random motion. The collision theory explains why reactions occur between these atoms, ions, and molecules. It also explains how it is possible to speed up or slow down reactions that are occurring.

The rate of reaction can be discussed in terms of three factors: collision frequency, the collision energy, and the geometric orientation. Collision frequency is the number of collisions per second. The collision frequency increases with the increase in temperature of the reaction.

When the temperature is increased, the average velocity of the particles is increased. The average kinetic energy of these particles is also increased. The result is that the particles will collide more frequently, because the particles move around faster and will encounter more reactant particles. Rate is increased just because the particles are, colliding more frequently. The major effect of increasing the temperature is that more of the particles that collide will have the amount of energy needed to have an effective collision. In other words, more particles will have the necessary activation energy.

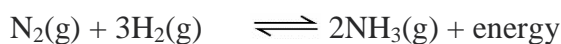
With an increase in temperature, there is an increase in energy that can be converted into activation energy in a collision.

Hence there is an increase in the number of effective collisions, which will increase the reaction rate

Q.4 a. Under what conditions synthesis of ammonia will give maximum yield by Haber's process. (2+2+2)

Answer:

If we look at the reaction of Ammonia;



It is an exothermic reaction where volume or number of moles are decreasing, so according to Le Chatelier's principle; the favorable conditions for ammonia production are:

1. High Pressure 200 atm (As number of moles are decreasing)
2. Low Temperature 400 °C (As the reaction is exothermic)
3. Removal of Product
4. Addition of reactants

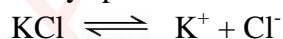
- 1) **By increasing pressure:** the yield can be maximized. Moles of product are less so by increasing pressure equilibrium will shift forward to increase the production of ammonia. 200 atm pressure is feasible.
- 2) **At low temperature:** The rate of a reaction increases as the temperature increases, so to increase the temperature until reactants cross the potential barrier. 400 °C is the temperature at which maximum yield is obtained. This is the optimum temperature.
- 3) **Removal of ammonia:** By removing the product, the equilibrium will shift forward and more product will be formed.
- 4) **Addition of reactants:** will also favour the formation of product. Increasing the concentration of reactants will shift the equilibrium forward making the conditions more favourable for the formation of more ammonia.

b. Describe dissolution of KCl and C₆H₁₂O₆ in water. (4+3)

Answer:

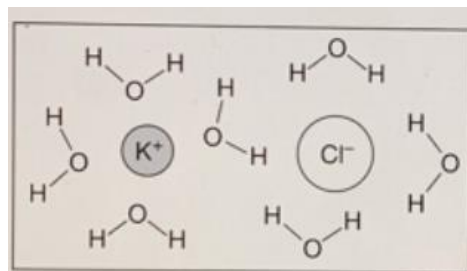
Dissolution of KCl:

When KCl is put in water, it immediately splits in to K⁺ and Cl⁻.



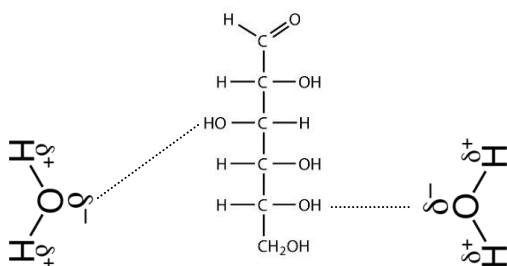
Water, H₂O, is a polar molecule, and has partial positive ⁺δ and negative charges ⁻δ. The hydrogens on water are partially positive, and the oxygen on water is partially negative.

Since opposite charges attract each other. So the oxygen of the water molecules are attracted to the K⁺, and the hydrogens are attracted to the Cl⁻. The water actually surrounds the K⁺ and Cl⁻ so that the opposite charges are allowed to be close together to form hydrate and process is called hydration (solvation). That is why KCl is dissolved in water.



Dissolution of C₆H₁₂O₆

C₆H₁₂O₆ is polar with OH group and water is also polar. Water, H₂O has partial positive ⁺δ and negative charges ⁻δ. The hydrogens on water are partially positive ⁺δ, and the oxygen on water is partially negative ⁻δ. Since opposite charges attract each other. So the oxygen of the water molecules are attracted to hydrogen of glucose to form hydrogen bonding. That is why glucose dissolve in water.

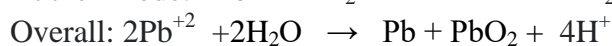


Dotted line show hydrogen bonding in between glucose and water.

Q.5 a. Demonstrate the reactions that occur in lead storage battery when it is recharged.

Answer:

The following half-cell reactions take place inside the cell during recharging:



During the charging process, the reactions at each electrode are reversed; the anode becomes the cathode and the cathode becomes the anode.

b. Sulphuric acid can be prepared by contact process using following reactions when one ton (1000kg) sulphur is used then how much SO_3 will be produced and how much oxygen is used? (3+4)



Answer:

Data:

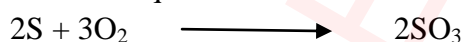
Mass of S = 1000kg = 1000000 g

First we calculate the moles of S.

Moles of S = $1000000 / 32$

Moles = 31250

Reaction equation:



Amount of SO_3 produced:

According to equation:

2 moles of S is producing : 2 moles of SO_3

1 moles of S is producing : $2/2$ moles of SO_3

31250 moles of S is producing : $2/2 \times 31250$ moles of SO_3

Mass of SO_3 = moles of $\text{SO}_3 \times$ molar mass of SO_3

Mass of SO_3 = 31250×80

Mass of SO_3 = 2500000 g = **2500 kg**

Amount of O_2 used:

According to the equation:

2 moles of S reacting with : 3 moles of O_2

1 moles of S reacting with : $\frac{3}{2}$ moles of O_2

31250 moles of S reacting with : $\frac{3}{2} \times 31250$ moles of O_2

31250 moles of S reacting with : 46875 moles of O_2

Mass of oxygen used = $46875 \times 32 = 1500000$ g = **1500 kg**

FBISE PAST PAPERS