



CHEMISTRY
HSSC-II
SECTION – A (Marks 17)

Time allowed: 25 Minutes

Section – A is compulsory. All parts of this section are to be answered on this page and handed over to the Centre Superintendent.

Deleting/overwriting is not allowed.

Do not use lead pencil.

حصہ اول لازمی ہے۔ اس کے جوابات اسی صفحہ پر دے کر نام مرکز کے حوالے کریں۔ کاٹ کر دوبارہ
کتنے کی اجازت نہیں ہے۔ اسے پینسل کا استعمال ممنوع ہے۔

Version No.			
4	0	9	1

ROLL NUMBER					

0	●	0	0
1	1	1	●
2	2	2	2
3	3	3	3
●	4	4	4
5	5	5	5
6	6	6	6
7	7	7	7
8	8	8	8
9	9	●	9

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6	6	6	6	6	6
7	7	7	7	7	7
8	8	8	8	8	8
9	9	9	9	9	9

Answer Sheet No. _____

ہر سوال کے سامنے دیے گئے، کرکولم کے مطابق درست دائرہ کو پر کریں۔
Invigilator Sign. _____

Fill the relevant bubble against each question according to curriculum:

Candidate Sign. _____

Question	A	B	C	D	A	B	C	D
1. Which of the following has highest ionization energy?	Lithium	Boron	Oxygen	Nitrogen	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
2. Which of the following oxides is acidic?	Na_2O	KO_2	MgO	NO_2	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
3. Pb^{+4} compounds are good:	Reducing agents	Oxidizing agents	Dehydrating agents	Hydrating agents	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
4. Carboxylic acids can be prepared by the reactions of Grignard reagent with:	Oxygen	Nitrogen	Carbon dioxide	Formaldehyde	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
5. To what category of molecules do enzymes belong?	Carbohydrates	Proteins	Lipids	Fats	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
6. Peroxyacetyl Nitrate (PAN) is an irritant to human beings and it affects:	Eyes	Nose	Stomach	Skin	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
7. Chemically nail polish remover is:	Acetone	Acetic acid	Ethanol	Ether	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
8. Which of the following electronic transitions requires highest amount of energy when a compound is interacted with ultraviolet radiations?	$n \rightarrow \pi^*$	$\pi \rightarrow \pi^*$	$\delta \rightarrow \delta^*$	$n \rightarrow \delta^*$	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
9. Height of peak in mass spectrum shows:	Number of isotopes	Mass number	Relative abundance of isotopes	Number of protons	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
10. Which of the following will NOT give iodoform test?	Formaldehyde	Acetaldehyde	Acetone	Ethanol	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
11. Oxidation of 2-Propanol gives:	Acetaldehyde	Formaldehyde	Acetone	Acetic acid	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
12. Nitration of phenol takes place at _____ than benzene.	Slower rate	Faster rate	Equal rate	Moderate rate	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
13. Reaction of alcohols with sodium metal gives:	Helium gas	Hydrogen gas	Carbon dioxide	Carbon monoxide	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
14. Geometrical isomerism is shown by:	Lactic acid	Acetic acid	1-Butene	2-Butene	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
15. Which of the following alkynes is acidic in nature?	1-Butyne	2-Butyne	2-Pentyne	3-Hexyne	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
16. Which type of functional group is present in compound $CH_3COOCOCH_3$?	Ester	Anhydride	Acid amide	Ether	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
17. The oxidation state of central metal (Mn) in the following complex $Na[Mn(CO)_5(F)_2]$ is:	+2	+3	-1	+1	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>

SUPPLEMENTARY TABLE

—2HA-I 2309—

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40



CHEMISTRY HSSC-II

Time allowed: 2:35 Hours

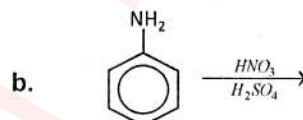
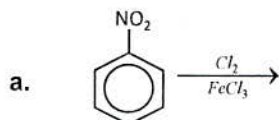
Total Marks Sections B and C: 68

NOTE: Answer any fourteen parts from Section 'B' and any two questions from Section 'C'. Write your answers neatly and legibly.

SECTION – B (Marks 42)

Q. 2 Attempt any FOURTEEN parts. All parts carry equal marks. (14 x 3 = 42)

- (i) How can the following conversions be carried out?
 - a. Nitro benzene → Aniline
 - b. Acetamide → Methyl amine
- (ii) Differentiate between 'Ethanol', '2-Propanol' and '2-Methyl 2-Propanol' on the basis of Lucas test.
- (iii) Write mechanism of the reaction when two molecules of Formaldehyde react with each other in the presence of $NaOH$.
- (iv) Write the preparation of Ethanol and 2-Propanol from Grignard reagent.
- (v) How is Nylon 6,6 prepared by condensation polymerization? Write its equations.
- (vi) Enlist the raw materials used for the preparation of lipstick.
- (vii) Briefly explain the trends in solubility of the hydroxides and sulphates of group II elements.
- (viii) Write any three advantages of atomic emission spectroscopy.
- (ix) Identify any three types of electronic transitions in Formaldehyde ($H-CHO$) when it is subjected to ultraviolet radiations
- (x) How can Ethanol and Acetone be identified by IR spectrum?
- (xi) Write chemical equations for the preparation of Acetone by hydration of a suitable alkyne.
- (xii) Briefly describe two causes that explain the inert nature of alkanes towards polar reagents.
- (xiii) Predict the major product in the following reactions



- (xiv) What is homologous series? Write its two characteristics.
- (xv) Write names of the following coordination compounds.
 - a. $Na_4[Mn(CN)_6]$
 - b. $Cr[(NO_2)_3(NH_3)_3]$
 - c. $[Ni(CO)_4(F)_2]SO_4$
- (xvi) Why Zn^{+2} complexes are colorless? Explain briefly with the help of d-d transitions phenomenon.
- (xvii) Why is ionization energy of nitrogen greater than that of oxygen? Although the ionization energy increases along the period.
- (xviii) Write chemical reactions to show the amphoteric nature of Al_2O_3 .
- (xix) Why K_2CO_3 is more stable than Li_2CO_3 ? Support your answer in terms of the polarizing ability of positive ion.
- (xx) Why is CCl_4 insoluble and $SiCl_4$ soluble in water? Explain your answer by writing mechanism for hydrolysis of group IV halides.

SECTION – C (Marks 26)

Note: Attempt any TWO questions. All questions carry equal marks. (2 x 13 = 26)

- Q. 3**
 - a. Write mechanisms for nitration and alkylation of Benzene.
 - b. What is Global warming? Write three adverse effects of each 'global warming' and 'acid rain' on environment.
- Q. 4**
 - a. What is esterification? Write mechanism for the acid catalyzed reaction between methanol and acetic acid.
 - b. Briefly explain the effect of:
 - i. Temperature
 - ii. pH and
 - iii. Substrate concentration
 on enzyme activity along with general graphical representation for each factor.
- Q. 5**
 - a. Differentiate beryllium (Be) from other members of its group in terms of its behaviors of:
 - i. Nitrides
 - ii. Carbides
 - iii. Oxides
 - iv. Reaction with alkalies
 - v. Reaction with Hydrogen
 - vi. Water of crystallization and its
 - vii. Hardness.
 - b. Explain three differences using 'Kinetic evidence', 'Stereochemical evidence' and 'Structure of substrate' between SN_2 and SN_1 reactions.

— 2HA-I 2307 —

SUPPLEMENTARY TABLE

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40



CHEMISTRY HSSC-II

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کھینے کی اجازت نہیں ہے۔ سبڈ پینسل کا استعمال ممنوع ہے۔

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Answer Sheet No. _____

ہر سوال کے سامنے دیے گئے، کریکولم کے مطابق درست دائرہ کو پر کریں۔ Invigilator Sign. _____

Fill the relevant bubble against each question according to curriculum: Candidate Sign. _____

Question	Candidate Sign.			
	A	B	C	D
1. Carboxylic acid reacts with metal to form salts with the evolution of:	CO_2	H_2	CO	CH_4
2. Reaction of Triglycerides with alkalis is called:	Hydration	Saponification	Hardening	Dehydration
3. Which of the following is correct formula of Ziegler catalyst?	$R_3Al + ZnCl_4$	$R_2Al + TiCl_4$	$R_3Al + TiCl_4$	$RAI + TiCl_4$
4. Oxidizing smog consists of high concentration of:	O_3	SO_2	NO_2	Cl_2
5. Which region of electromagnetic spectrum is involved in bonds vibrations?	Ultraviolet	Infra red	Radio waves	Microwaves
6. The characteristic signal of Carbonyl Carbon of acetone in I.R spectrum will be detected at:	$3200cm^{-1}$	$1200cm^{-1}$	$800cm^{-1}$	$1720cm^{-1}$
7. Which of the following aldehydes will be most reactive towards Nucleophiles?	Benzaldehyde	Acetaldehyde	Formaldehyde	Propanaldehyde
8. Which of the following has highest ionization energy?	Al	Si	P	S
9. Which of the following alcohols will produce least stable alkoxide ion when $O-H$ bond breaks?	Methanol	Ethanol	2-Propanol	2-Methyl 2-Propanol
10. 2,3-dimethyl-2-butene on reaction with O_3/H_2O yields two molecules of:	Acetaldehyde	Acetone	Formaldehyde	Benzaldehyde
11. Which of the following alcohols cannot be synthesized when Grignard reagent reacts with aldehydes and ketones?	Ethanol	Isobutyl alcohol	Propanol	Methanol
12. No. of chiral centres in lactic acid $COOH(CH)OH$ are: $\begin{array}{c} \\ CH_3 \end{array}$	Two	One	Three	Four
13. The general formula of cycloalkanes resembles with.	Alkenes	Alkynes	Alkanes	Benzene
14. The only alkaline earth metal which releases H_2 gas when reacted with alkalis is:	Be	Mg	Ca	Sr
15. Which of the following is most Paramagnetic?	Fe^{3+}	Sc^{3+}	Ti^{4+}	V^{4+}
16. Coordination sphere in $[Cu(en)_3]SO_4$ is:	Neutral	Positively charged	Negatively Charged	Amphoteric
17. Which of the following hydroxides of group II elements is least soluble in water?	$Ca(OH)_2$	$Mg(OH)_2$	$Be(OH)_2$	$Sr(OH)_2$

SUPPLEMENTARY TABLE

—2HA-I 2309 HA—

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CHEMISTRY HSSC-II

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Total Marks Sections B and C: 68

NOTE: Answer any fourteen parts from Section 'B' and any two questions from Section 'C'. Write your answers neatly and legibly.

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Q. 2 Attempt any FOURTEEN parts. All parts carry equal marks. (14 x 3 = 42)

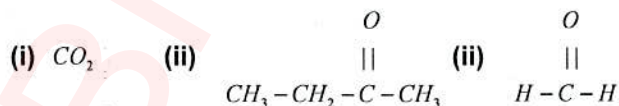
- (i) Why group II carbonates are thermally unstable as compared to group I carbonates?
- (ii) Why salts of Be^{+2} ion cannot have more than four molecules of water of crystallization?
- (iii) Write chemical reactions to show the amphoteric nature of BeO .
- (iv) Why CO_2 is gas while SiO_2 is solid at room temperature? Draw structural diagrams of CO_2 and SiO_2 .
- (v) Describe with justification of the anomalous trends in the ionization potential of 3rd period elements.
- (vi) Write IUPAC names of following coordination compounds.
 - a. $(NH_4)_2[Pt(Cl)_6]$
 - b. $[CO(F)_6(Br)_2]SO_4$
 - c. $[Ni(CO)_5]$
- (vii) How V_2O_5 acts as a catalyst in contact process in the given reaction? $SO_2 + \frac{1}{2}O_2 \xrightarrow{V_2O_5} SO_3$
- (viii) What is meant by a. *Partial Synthesis* b. *Total Synthesis* of organic compounds in lab?
- (ix) Why Pb^{+4} compounds are less stable and act as an oxidizing agent? Explain with help of inert pair effect.
- (x) Why bond enthalpy of $F-F$ is less than $Cl-Cl$? Although size of Chlorine is bigger than that of Fluorine.
- (xi) Differentiate between metamerism and tautomerism with a suitable example of each.
- (xii) Which two factors determine the reactivity of alkyl halides? Write correct order of reactivity.
- (xiii) Write product and mechanism for the following reaction. $CH_2 = CH_2 + Br_2 \rightarrow ?$
- (xiv) Differentiate between 1-butyne and 2-butyne by ozonolysis.
- (xv) Which of the following is more acidic? Give reason. a. 2-Methyl Phenol b. 2-Nitro Phenol
- (xvi) How can Ethanol and Methanol be differentiated by iodoform test? Write chemical equation for the test.
- (xvii) Briefly explain any three sources by which raw materials are obtained for chemical industries.
- (xviii) What is spectroscopy? Write its principle.
- (xix) Write three applications of atomic absorption spectroscopy.
- (xx) Differentiate between qualitative and quantitative analysis with suitable example of each.

SECTION – C (Marks 26)

Note: Attempt any TWO questions. All questions carry equal marks. (2 x 13 = 26)

- Q. 3 a. What are carbohydrates? Briefly describe
i. Monosaccharides, ii. Disaccharides and iii. Polysaccharides with suitable examples.
- b. Write mechanism for the formation of an ester by a reaction between ethanol and acetic acid catalyzed by an acid.

- Q. 4 a. What is sulphonation? Write mechanism for sulphonation of Benzene.
- b. Write reactions and predict the products when Methyl Magnesium Bromide (CH_3MgBr) reacts with:



- Q. 5 a. What is acid rain? How it occurs? Write its four adverse effects on environment.
- b. Justify the given order of acidity with reason. *Carboxylic Acid > Phenol > Alcohol*

— 2HA-I 2309 HA —

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